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Hydrogen Bonding of O-Ethylxanthate Compounds and Neutron Structural Determination of C-H•••S Interactions

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The hydrogen bonding in mixed sulfur/oxygen acceptor systems can be thoroughly investigated using the Oethylxanthate (or O-alkyldithiocarbonate) family of anions. A series of O-ethylxanthate salts (guanidinium, methylammonium, dimethylammonium, trimethylammonium, tetramethylammonium, tetraethylammonium, and tetrapropylammonium), were structurally characterised using synchrotron X-ray spectroscopy to demonstrate the influence of the cation on the overall packing of the salts into either 3-D, 2-D or 1-D hydrogenbonded arrangements. The protic cations vary in the number of available hydrogen bond donors which in turn affects the dimensionality of the hydrogen-bonded networks that form. The guanidinium cation gives rise to a 3-D hydrogen-bonded network due to the large number of NH hydrogen bond donors, whereas the use of ammonium cations, of the nature MexH4-xN+ (x = 1 - 3), demonstrates the decreasing dimensionality of the structures as the number of hydrogen bond donors decreases. Aprotic cations were also studied to demonstrate the influence of C-H...S interactions on the overall packing arrangement of the structure, in the absence of strong hydrogen bond donor groups. Laue neutron diffraction data was used to locate weak C-H...S hydrogen bonds in (Me4N)(EtXn) through the location of the exact positions of the CH hydrogen donors. The neutron data demonstrates the unequivocal presence of CH+++S hydrogen bonding, with the H+++S distance significantly shorter than the sum of the van der Waals' radii (shortest interaction 2.67 Å compared with 3.00 Å).

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Summary

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